Complete the table by writing names and formulas for the compounds formed when the cations combine with the anions. If there is no charge for an ion, you should place it.

|  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- |
| Ions | Fe+2 | Al \_\_ | Na \_ | Ca \_\_ | NH4\_\_ | Fe+3 | Ni +2 | Hg +2 |
| CO3 \_**2-** | **FeCO3**  **Iron (II) Carbonate** |  |  |  |  |  |  |  |
| O \_\_ |  |  |  |  |  |  |  |  |
| C2H3O2\_\_ |  |  |  |  |  |  |  |  |
| OH \_\_ |  |  |  |  |  |  |  |  |
| HCO3 \_\_ |  |  |  |  |  |  |  |  |
| PO4 \_\_ |  |  |  |  |  |  |  |  |
| SO3 \_\_ |  |  |  |  |  |  |  |  |
| ClO3 \_\_ |  |  |  |  |  |  |  |  |
| SO4 \_\_ |  |  |  |  |  |  |  |  |
| Cl \_\_ |  |  |  |  |  |  |  |  |

Complete the table by writing the formulas for the compounds formed when the anions combine with the cations. Write the formula and charge of each ion below its name.

|  |  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- | --- |
| Ions | bromide  **Br-** | Hydrogen carbonate  \_\_\_\_\_\_ | Sulfide  \_\_\_\_\_\_ | Acetate  \_\_\_\_\_\_ | Sulfate  \_\_\_\_\_\_\_ | Phosphate  \_\_\_\_\_\_\_\_ | Oxide  \_\_\_\_\_\_\_ | Chlorate  \_\_\_\_\_\_\_ | Nitride  \_\_\_\_\_\_ |
| Calcium  **Ca2+** | **CaBr2** |  |  |  |  |  |  |  |  |
| Strontium  \_\_\_\_\_\_\_\_ |  |  |  |  |  |  |  |  |  |
| Ammonium  \_\_\_\_\_\_\_\_ |  |  |  |  |  |  |  |  |  |
| Aluminum  \_\_\_\_\_\_\_\_ |  |  |  |  |  |  |  |  |  |
| Iron (III)  \_\_\_\_\_\_\_\_ |  |  |  |  |  |  |  |  |  |
| Lead (IV)  \_\_\_\_\_\_\_ |  |  |  |  |  |  |  |  |  |
| Silver (I)  \_\_\_\_\_\_\_\_ |  |  |  |  |  |  |  |  |  |
| Gold (III)  \_\_\_\_\_\_\_\_ |  |  |  |  |  |  |  |  |  |
| Potassium  \_\_\_\_\_\_\_\_ |  |  |  |  |  |  |  |  |  |
| Mercury(I)  \_\_\_\_\_\_\_\_ |  |  |  |  |  |  |  |  |  |
| Barium  \_\_\_\_\_\_\_\_ |  |  |  |  |  |  |  |  |  |